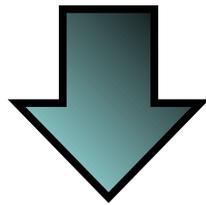


#13 Stoichiometry: Determination of an unknown

Use the same balance for all weighing

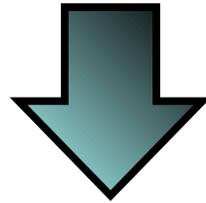
First, you react citric acid with pure sodium bicarbonate and measure the volume of CO_2 gas formed:

citric acid + pure sodium bicarbonate



CO_2 + other products

Second, you react citric acid with impure sodium bicarbonate and measure the volume of CO_2 gas formed:
citric acid + impure sodium bicarbonate



CO_2 + other products

If the unknown was pure sodium bicarbonate, you would expect same volume of gas to form, right?

But the unknown is NOT pure sodium bicarbonate. By comparing volumes of gas produced, you can determine the % sodium bicarbonate in the impure mixture.

Set up a ratio like this:

$$\frac{\text{mass pure sodium bicarbonate}}{\text{mL CO}_2 \text{ formed}} = \frac{\text{mass impure sodium bicarbonate}}{\text{mL CO}_2 \text{ formed}}$$

Once you have the mass of sodium bicarbonate in the impure sample, you can calculate the % of sodium bicarbonate in the impure sample.....

= % SODIUM BICARBONATE
IN IMPURE SAMPLE

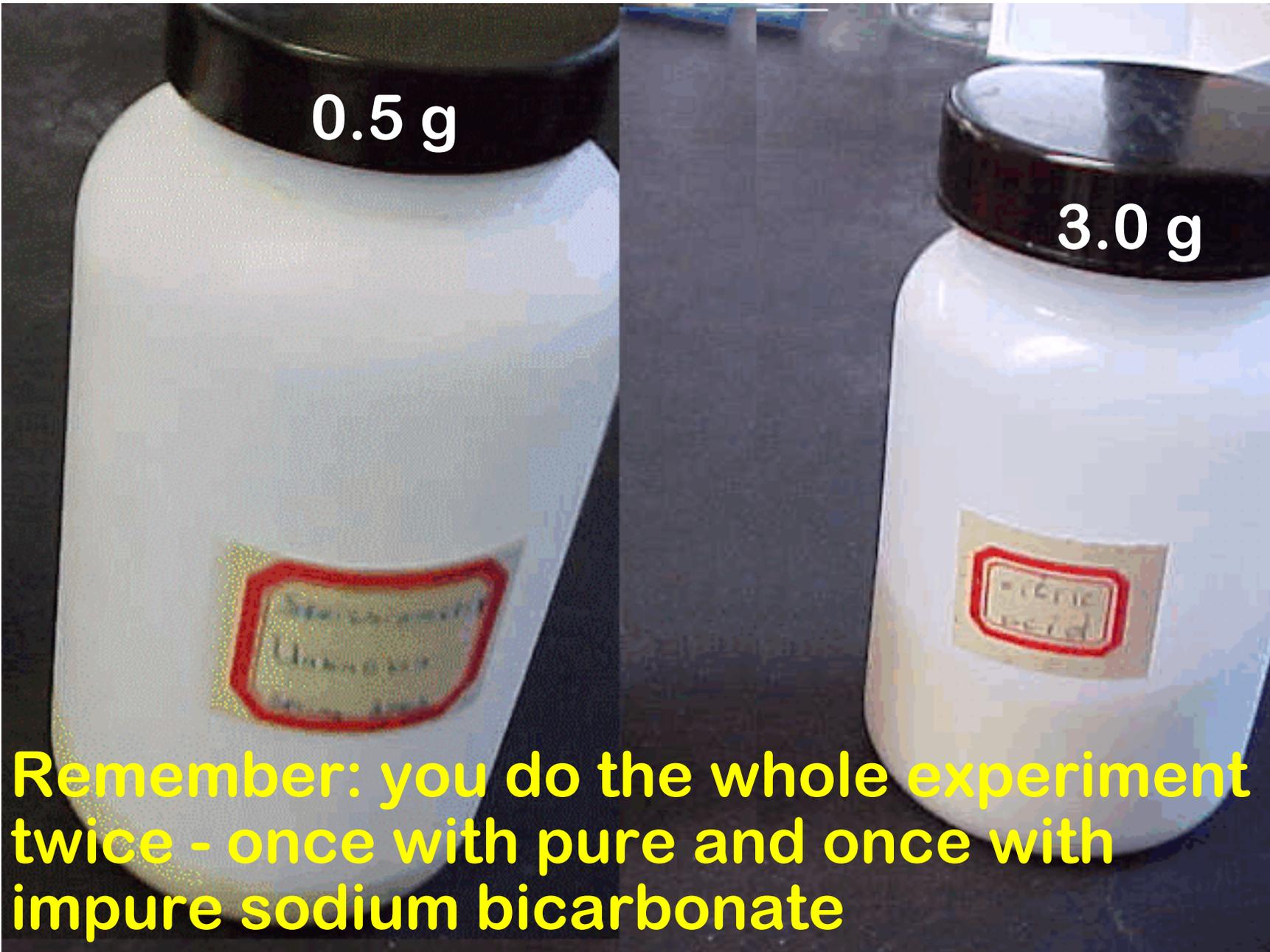
$$\frac{\text{mass of bicarbonate in impure sample}}{\text{mass of unknown sample (= } \sim 0.5 \text{ g)}} \times 100$$



0.5 g

3.0 g

NOTE: be careful weigh out the samples - don't mix up the masses

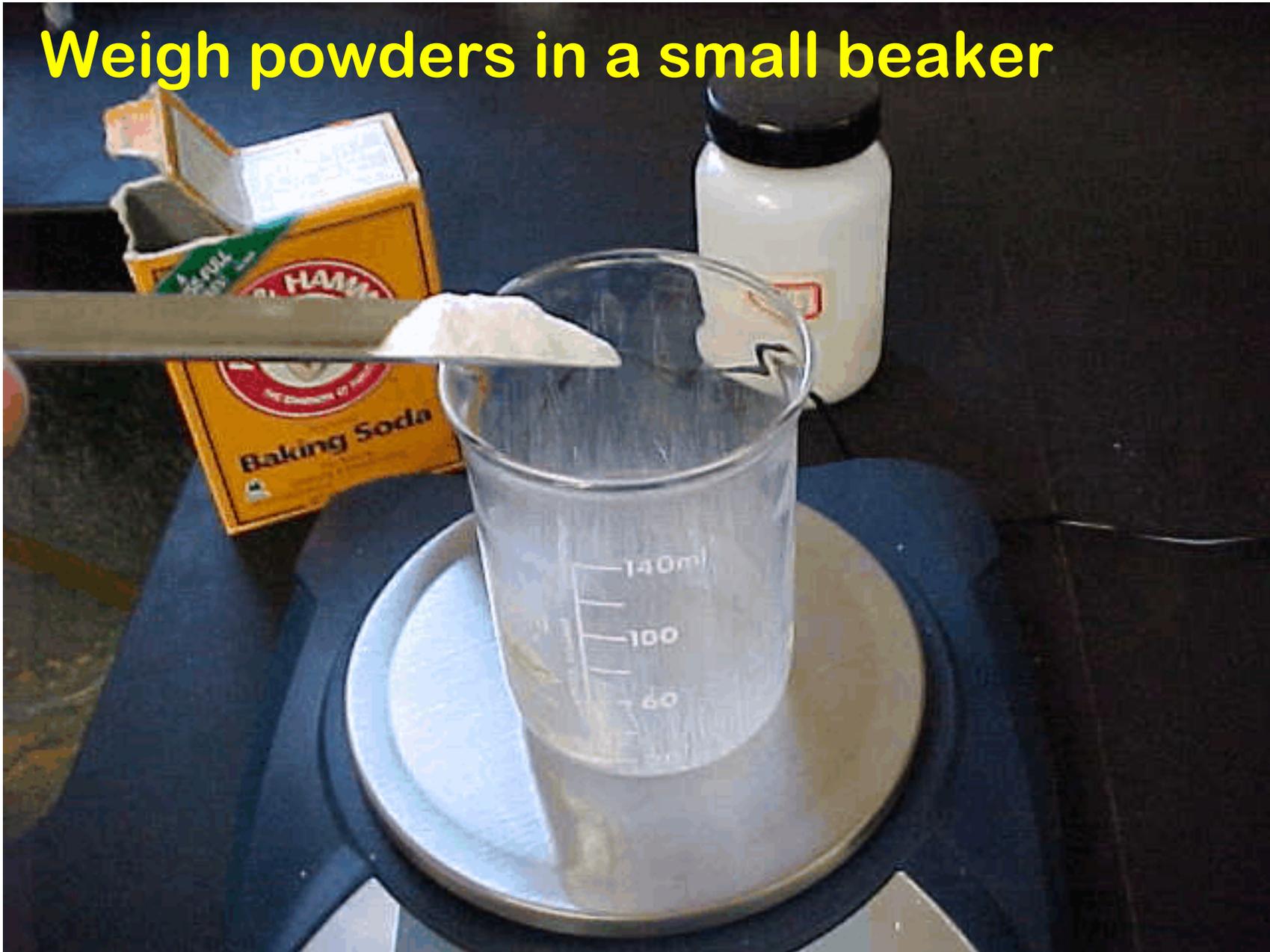


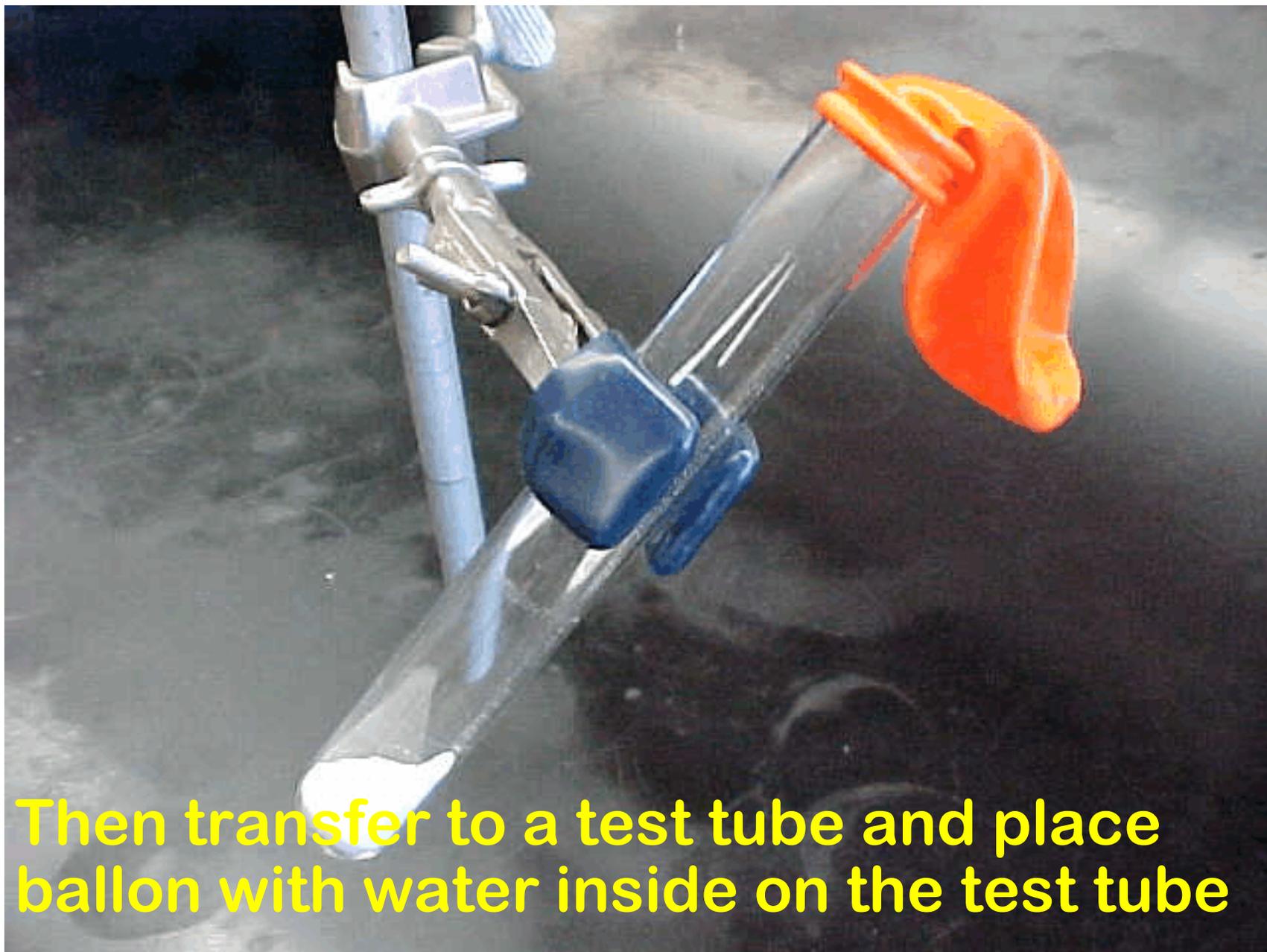
0.5 g

3.0 g

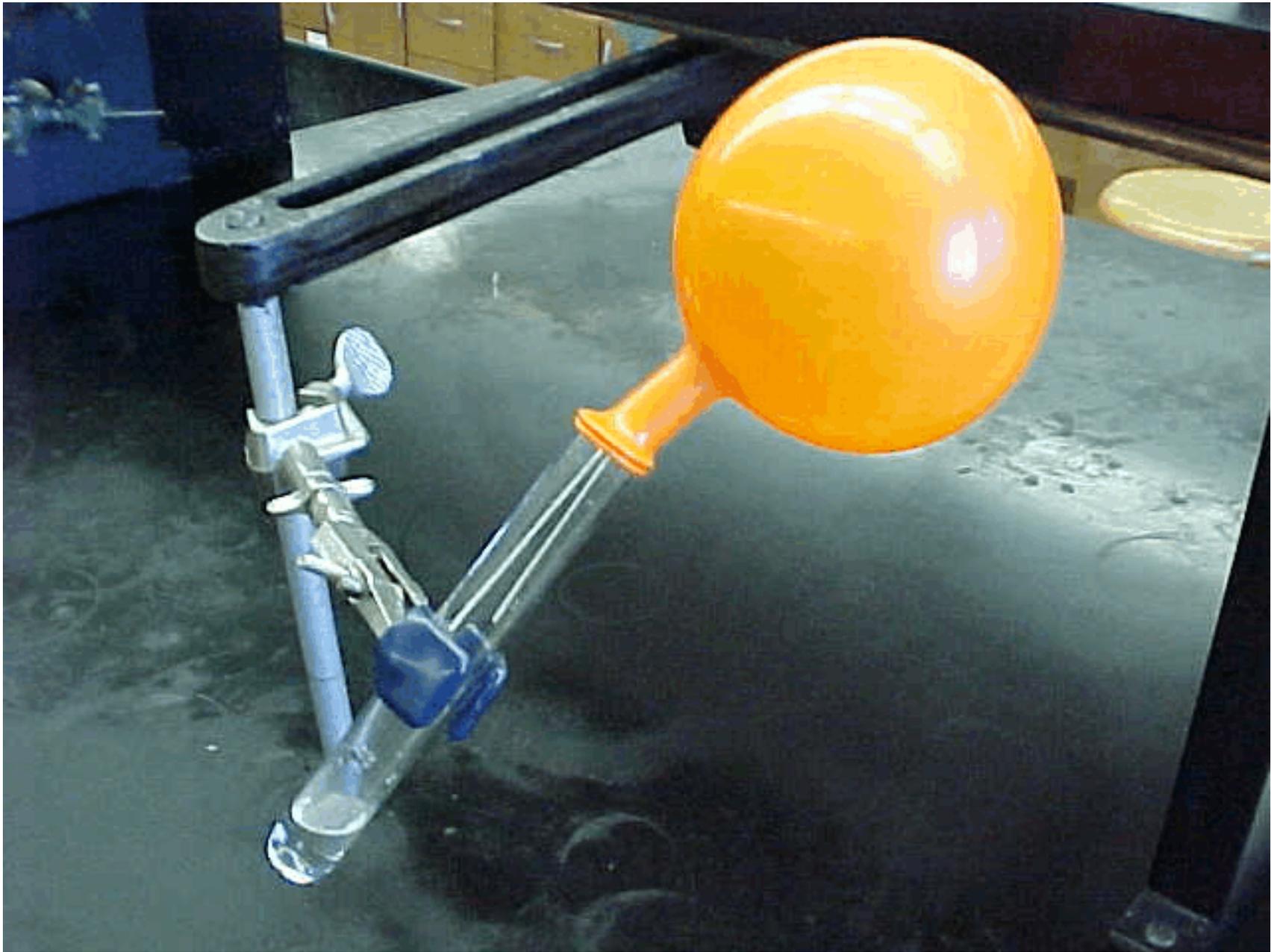
Remember: you do the whole experiment twice - once with pure and once with impure sodium bicarbonate

Weigh powders in a small beaker

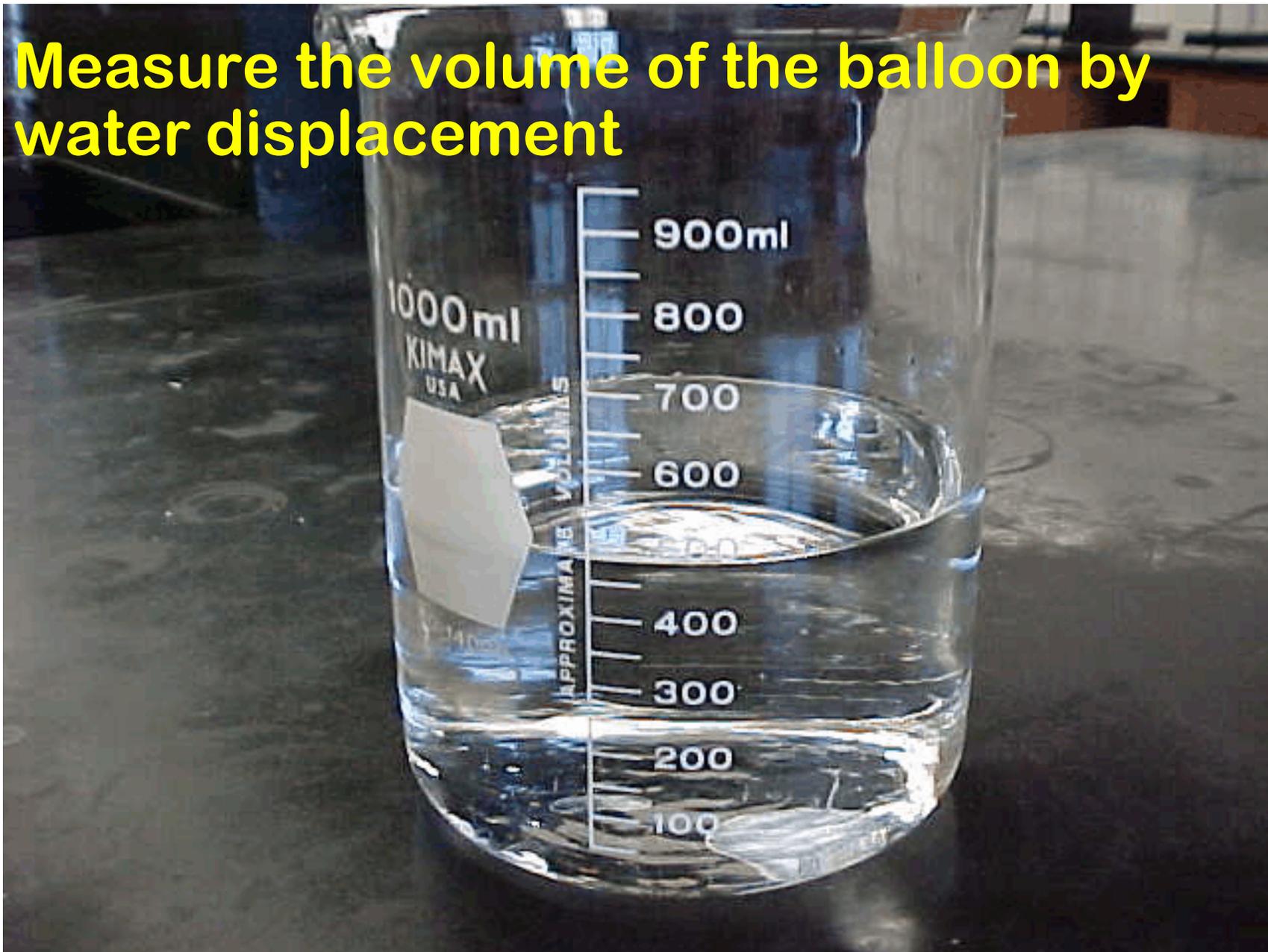




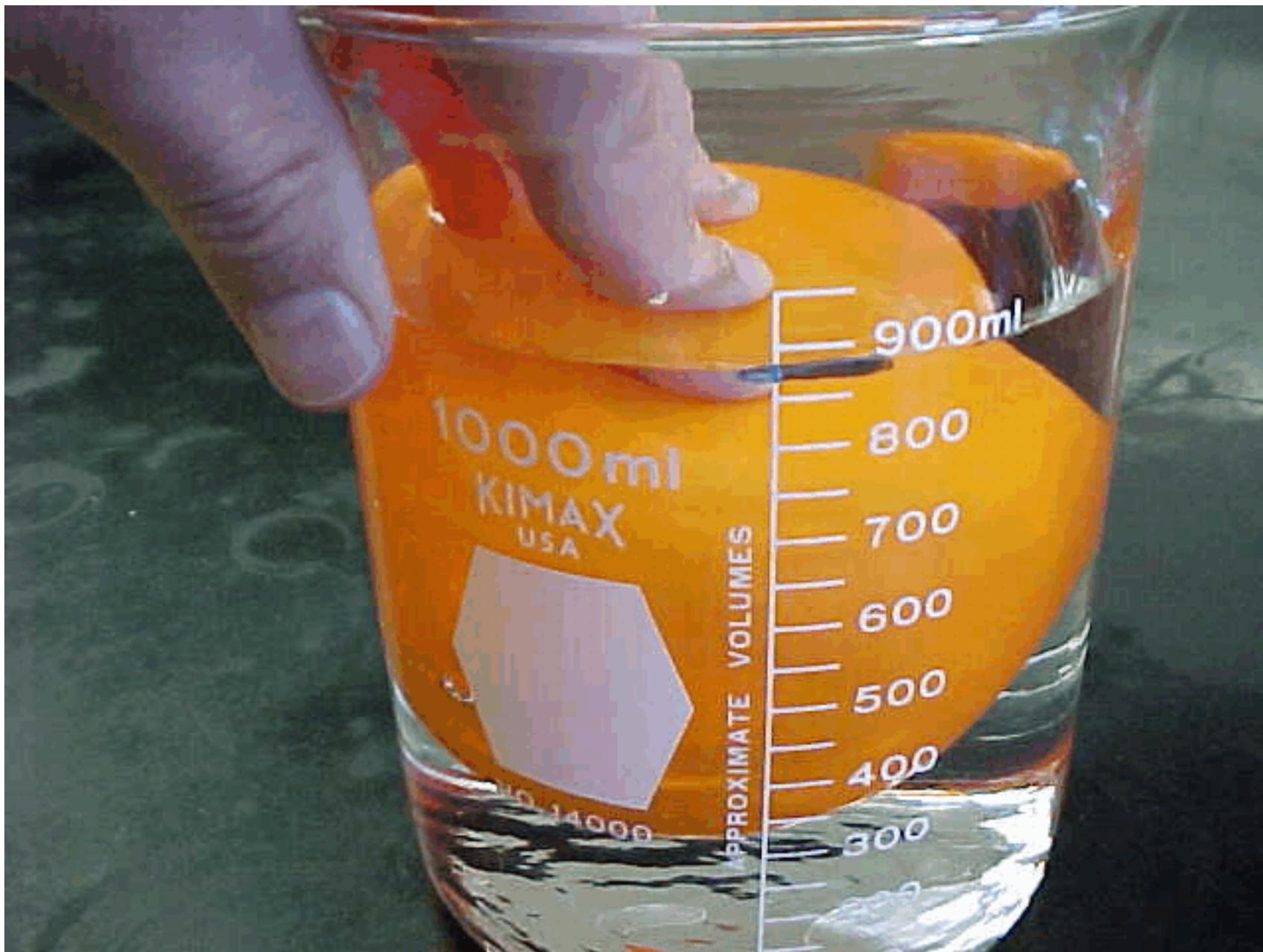
**Then transfer to a test tube and place
ballon with water inside on the test tube**

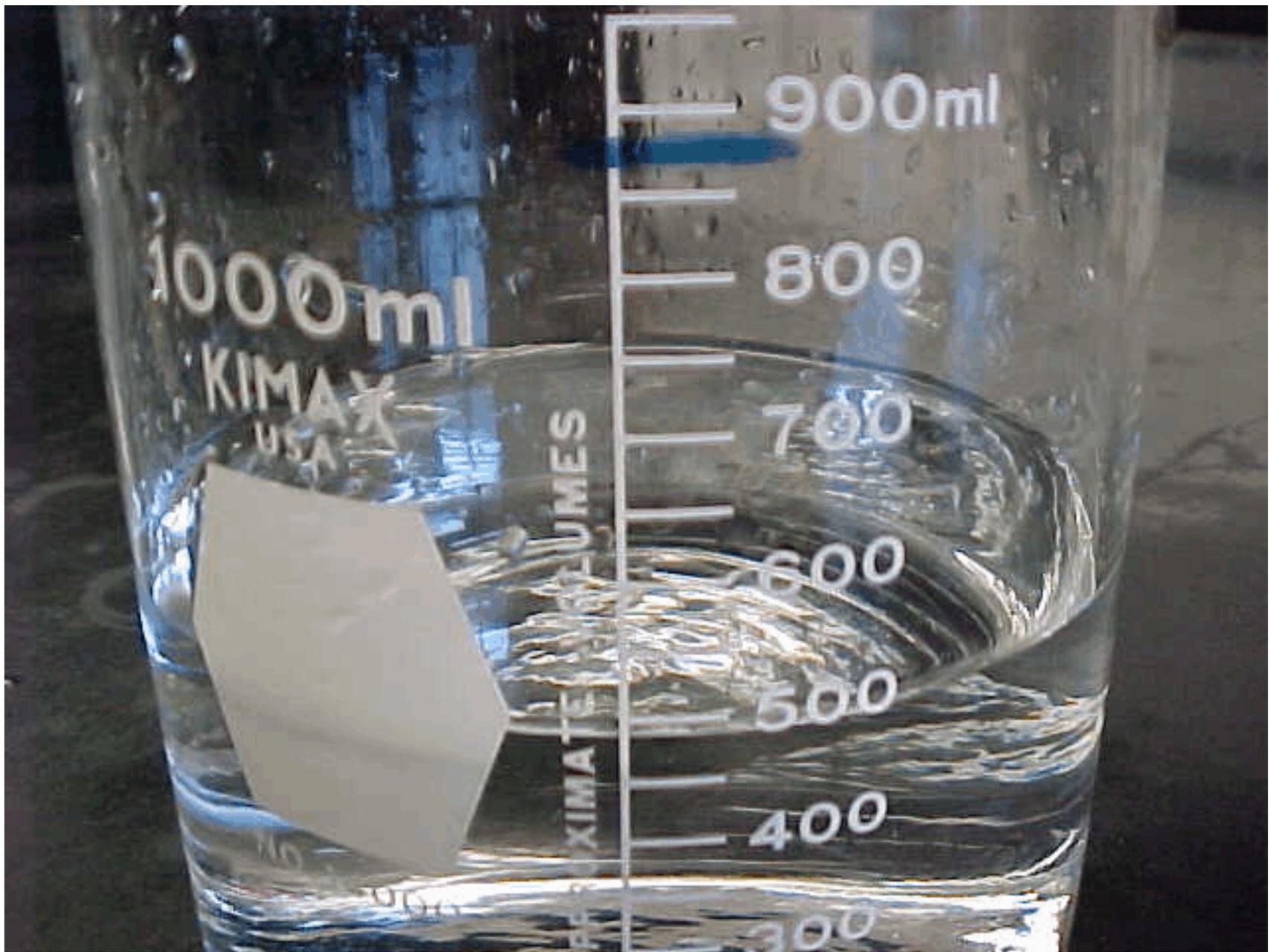


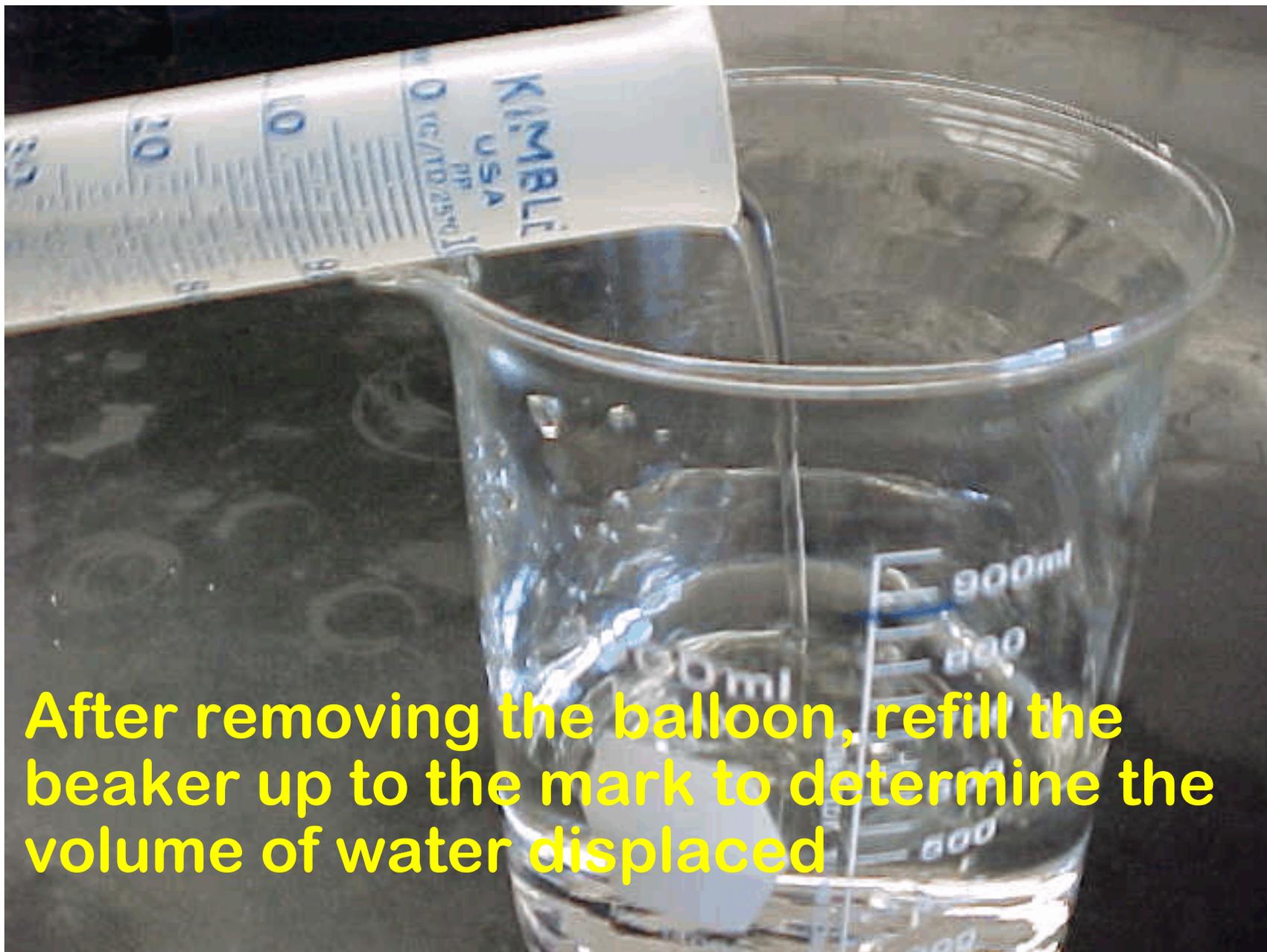
Measure the volume of the balloon by water displacement











After removing the balloon, refill the beaker up to the mark to determine the volume of water displaced

