

Preparation & Analysis of Alum

This is a 2-week experiment.

First week: Exp 32: Part A

You will prepare alum from scrap aluminum this week, then analyse it for purity next week

Write up part A and part B separately. And Don't forget, each person hands in their own lab report

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Notes

Follow lab instructions under “experimental procedure.”

Always use the same balance for weighing.

Wear safety goggles at all times in the lab.

Alum is $\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}$

remember, this means there are 12 water molecules to include in the molecular mass of alum

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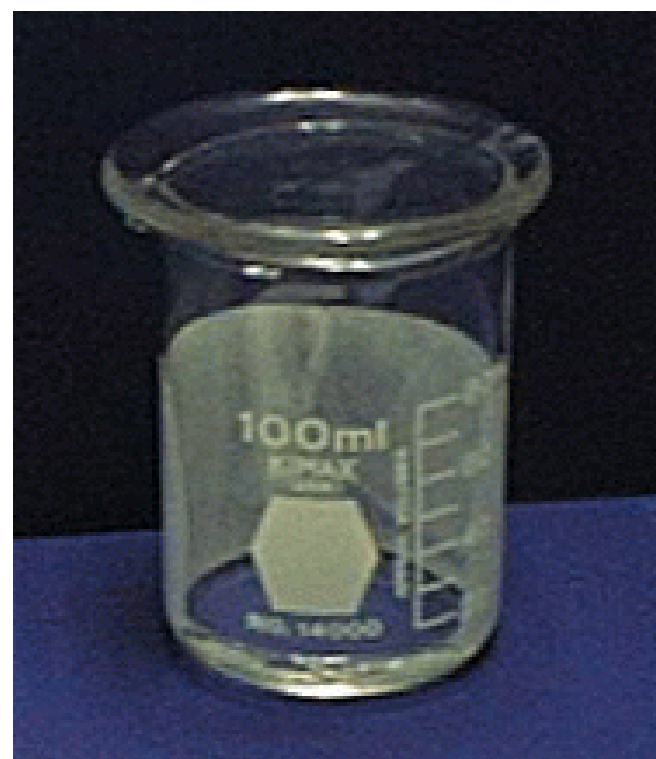
Notes

The reaction generates hydrogen gas - no flames

Place watchglass over beaker during heating (curved side of watchglass down).

This prevents solution boiling dry - see photo.

Boil gently!



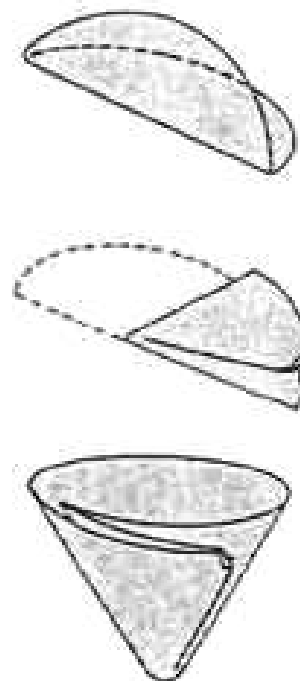
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Notes

When filtering, fold filterpaper in half, then half again.

Open to make a cone that sits in funnel
- see diagram

If alum does not crysallize, scratch the inside bottom of the beaker with a glass rod while cooling.



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Notes

Leave the moist alum to dry in your drawers until next lab.

Lab cannot be written up until the mass of dry alum determined next week.

Write up part A and part B separately.