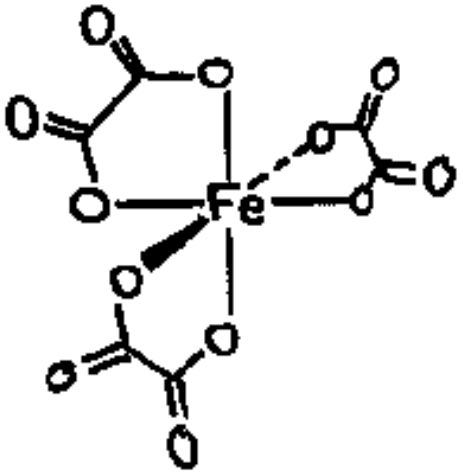
#33 Prep of potassium trisoxalatoferrate(III)

Part A

NOTE: This is a 2-week experiment. First week (Part A) you will prepare the compound, second week (Part B) you will analyze it.

The green product you are preparing is a coordination complex - details will be discussed in lectures.



The complex has the structure & formula: $K_3[Fe(C_2O_4)_3].3H_2O$

Don't forget the 3 waters of hydration in the molecular mass calculation

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Follow directions in the lab manual for the preparation. Very straight forward.

After getting the yellow precipitate, allow the solid to settle for a few minutes and carefully decant the liquid off, leaving the solid behind. If you lose a little of the solid, that's okay.

In the final step, after adding the hydrogen peroxide, heating, and filtering, the solution should look a lime-green color.

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It days some time for the green product to crystallize. You will collect the product next week, weigh it, and analyze it.

Both parts of the experiment will be due week after next. Part A should contain answers to the six questions (page 33-5); Part B has three questions (same page). Both parts can be handed in together.

When you isolate the product next week, it should look something like this 6