

#22 Kinetics

Objective to find x and y in rate law:

$$\text{Rate} = k[\text{I}^-]^x[\text{H}_2\text{O}_2]^y$$

Note: all solutions are prepared, but show calculations for Q1 on p 22-2 in your lab report

Do part A only

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Run

Follow instructions to mix the 5 solutions in the 2O_2 to start the reaction.

Run #3

Follow instructions to mix the 5 solutions in the 2O_2 to start the reaction.

Take

in correct order

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Compare
the KI concentration affects the rate.

Compare
the H₂O₂ concentration affects the rate.

Need watch with second hand